Summer Assignment Problem Set

1.66 Suppose you try to define your own temperature scale using the freezing point (-11.5◦C) and boiling point

 (197.6◦C) of ethylene glycol. If you set the freezing point as 0◦G and the boiling point as 100◦G, what is

 the freezing point of water on this new scale?

1.67 The liquid substances mercury (density = 13.5 g/mL) water (1.00g/mL), and cyclohexane (0.778g/mL) do

 not form a solution when mixed, but separate in distinct layers. Sketch how the liquids would position

 themselves in a test tube.

1.72 Automobile batteries contain sulfuric acid, which is commonly referred to as “battery acid.” Calculate

 the number of grams of sulfuric acid in 0.500 L of battery acid if the solution has a density of 1.28 g/mL

 and is 38.1% sulfuric acid by mass.

1.75 A package of aluminum foil contains 50 ft2 of foil, which weighs approximately 8.0 oz. Aluminum has a

 density of 2.70 g/cm3. What is the approximate thickness of the foil in millimeters?

2.87 Identify the element represented by each of the following symbols and give the number of protons and

74

33

127

53

 neutrons in each a) X b) X

2.90 The element lead (Pb) consists of four naturally occurring isotopes with atomic mass 203.97302,

 205.97444, 206.97587, and 207.97663 amu. The relative abundances of these four isotopes are 1.4, 24.1,

 22.1, and 52.4%, respectively. From these data, calculate the atomic weight of lead.

2.98 Name each of the following oxides. Assuming that the compounds are ionic, what charge is associated

 with the metallic element in each case? a. NiO b. MnO2 c. Cr2O3 d. MoO3

2.102 Give the chemical names of each of the following familiar compounds: a. NaCl (table salt)

 b. NaHCO3 (baking soda) c. NaOCl (in many bleaches) d. NaOH (caustic soda)

 e. (NH4)2CO3 (smelling salts) f. CaSO4 (plaster of Paris)

2.104 Many ions and compounds have very similar names, and there is great potential for confusing them.

 Write the correct chemical formulas to distinguish between

 a. calcium sulfide and calcium hydrogen sulfide b. hydrobromic acid and bromic acid

 c. aluminum nitride and aluminum nitrite d. iron(II) oxide and iron(III) oxide

 e. ammonia and ammonium ion f. potassium sulfite and potassium bisulfite

 g. mercurous chloride and mercuric chloride h. chloric acid and perchloric acid.

2.105 The compound cyclohexane is an alkane in which six carbon atoms form a ring. The partial structural

 formula of the compound is as follows:

 a. Complete the structural formula for cyclohexane.



 b. Is the molecular formula for cyclohexane the same as

 that for *n*-hexane, in which the carbon atoms are in a

 straight line? If possible, comment on the source of

 any differences.

3.82 The effectiveness of nitrogen fertilizers depends on both their ability to deliver nitrogen to plants and the

 amount of nitrogen they can deliver. Four common nitrogen-containing fertilizers are ammonia,

 ammonium nitrate, ammonium sulfate, and urea [(NH3)2CO]. Rank these fertilizers in terms of the mass

 percentage nitrogen they contain.

3.87 The koala dines exclusively on eucalyptus leaves. Its digestive system detoxifies the eucalyptus oil, a

 poison to other animals. The chief constituent in eucalyptus oil is a substance called eucalyptol, which

 contains 77.87% C, 11.76% H, and the remainder O. a. What is the empirical formula for this substance?

3.99 When a mixture of 10.0 g of acetylene (C2H2) and 10.0 g of oxygen (O2) is ignited, the resultant

 combustion reaction produces CO2 and H2O. a. Write the balanced chemical equation for this reaction.

 b. Which is the limiting reactant? c. How many grams of C2H2 are present after the reaction is

 complete?

4.92 The reaction between Cd(NO3)2 and Na2S produces a precipitate. What is the identity of the precipitate?

 What ions remain in solution? Write the net ionic equation for the reaction.

5.98 A coffee-cup calorimeter contains 150.0 g of water at 25.1◦C. A 121.0-g block of copper metal is heated

 to 100.4◦C by putting it in a beaker of boiling water. The specific heat of Cu(s) is 0.385 J/g-K. The Cu is

 added to the calorimeter, and after a time the contents of the cup reach a constant temperature of 30.1◦C.

 a. Determine the amount of heat, in J, lost by the copper block.

6.15 a. What is the frequency of radiation that has a wavelength of 10 m, about the size of a bacterium?

 b. What is the wavelength of radiation that has a frequency of 5.50x1014 s-1? c. Would the radiations in

 part “a” or part “b” of this problem be visible to the human eye?

 

6.73 What is wrong with the following electron configurations for atoms in their ground states? a. 1s22s23s1

 b. [Ne]2s22p3 c. [Ne]3s23d5

7.45 Based on their positions in the periodic table, predict which atom of the following pairs will have the

 larger first ionization energy: a. Cl, Ar b. Be, Ca c. K, Co d. S, Ge e. Sn, Te

7.48 Write the electron configurations for the following ions, and determine which have noble-gas

 configurations: a. Cr3+ b. N3-

8.88 Calculate the formal charge on the indicated atom in each of the following molecules or ions: a. the

 central oxygen atom in O3

9.79 From their Lewis structures, determine the number of  and  bonds in each of the following molecules

 or ions: b. thiocyanate ion, NCS- c. formaldehyde, H2CO

10.91 A gas bubble with a volume of 1.0 mm3 originates at the bottom of a lake where the pressure is 3.0 atm.

 Calculate its volume when the bubble reaches the surface of the lake where the pressure is 695 torr,

 assuming that the temperature doesn’t change.

10.93 To minimize the rate of evaporation of the tungsten filament, 1.4x10-5 mol of argon is placed in a 600-cm3

 lightbulb. What is the pressure of argon in the lightbulb at 23◦C?

11.56 The normal melting and boiling points of O2 are -218◦C and -183◦C respectively. Its triple point is at

 -219◦C and 1.14 torr, and its critical point is at -119◦C and 49.8 atm. a. Sketch the phase diagram for O2,

 showing the four points given and indicating the area in which each phase is stable.

13.24 Referring to the following diagram,



 determine the mass of each of the

 following salts required to form a

 saturated solution in 250 g of water

 at 30◦C: a. KClO3

13.89 Fish need at least 4 ppm dissolved O2 for survival. A. What is this concentration in mol/L?

13.69 Using data from the following table, calculate the freezing and boiling points of each of the following

 solutions: a. 0.22 *m* glycerol (C3H8O3) in ethanol.

 

14.7 The following diagram shows the reaction profile of a reaction. Label the components indicated by the

 boxes.

 

15.14 Write the expressions for Keq for the following reactions. In each case indicate whether the reaction is

 homogeneous or heterogeneous. a. 2O3(g) ⇆ 3O2(g) b. Ti(s) + 2Cl2(g) ⇆ TiCl4(l)

16.39 Complete the following table by calculating the missing entries and indicating whether the solution is

 acidic or basic.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| [H+] | [OH-] | pH | pOH | acidic or basic |
| 7.5x10-3 M |  |  |  |  |

17.6 Match the following descriptions of titration curves with the diagrams:

 a. strong acid added to strong base c. strong based added to strong acid

 b. strong based added to weak acid d. strong base added to polyprotic acid

 

 (i) (ii) (iii) (iv)